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number of electrons present in an atom of this element is 8 What is the maximum number of electrons that can occupy the second energy level? Chapter 4 Atomic Structure Test Review Flashcards | Quizlet The sum of the protons and neutrons in the nucleus of an atom. $\text{MASS\#} = \text{\# protons (atomic number)} + \text{\# neutrons}$. $\text{\# NEUTRONS} = \text{Mass \#} - \text{\# Protons (atomic number)}$ $\text{\# Protons} = \text{Mass \#} - \text{\# Neutrons}$. ISOTOPES: Are atoms of the same element that have different numbers of neutrons and different mass numbers. Chapter 4: Atomic Structure Notes Flashcards | Quizlet Chemistry (12th Edition) answers to Chapter 4 - Atomic Structure - 4 Assessment - Page 122

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in atomic structure..
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the question..Chapter 4
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structure of the atom.
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paper in half
lengthwise. Make the
back edge about 2 cm
longer than the front
edge. Step 2 Fold into
thirds. Step 3 Unfold

and cut along one fold
line to make one small
tab and one large tab.
Step 4 Label as shown.
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Foldable with Section
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The Following

Questions Circle The Best Answer

1 Which Of The Following Are Isotopes Of Each Other

A ^{14}C And ^{14}N B ^3H And ^4He C ^2H And ^1H D None Of These

2 The Net Charge On An Atom That Has 13 Protons 12 Neutrons And 10 Electrons Is

A 5 B 3 C 3 D 9

3 Particles That Make Up Atoms Are Called Subatomic Particles And Are

Chapter 4 Atomic Structure Review Worksheet

Answers...Chapter 4 Atomic Structure Answers

(i) Atomic number = number of protons + number of electrons

(ii) Mass number = number of protons + number of neutrons

(iii) Atomic mass = number of protons = number of neutrons

(iv) Atomic number = number of protons = number of

electrons (a) (i) and (ii)
 (b) (i) and (iii) (c) (ii)
 and (iii) (d) (ii) and (iv)
 Answer. Answer: (d) (ii)
 and (iv) MCQ Questions
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 atoms be seen
 answer the challenges
 to his ideas,
 Democritus's atomic
 theory was eventu-
 ment, a narrow beam
 of alpha particles was
 aimed at a thin sheet
 of gold foil. Chapter 4
 Atomic Structure
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 Structure Answer Key -
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 Answer: (i) Volume of
 the sphere =
 $\frac{4}{3}\pi R^3$ Let
 R be the radius of the
 atom and r be that of
 the nucleus. Question
 6. Show
 diagrammatically the
 electron distribution in
 a sodium atom and a
 sodium ion and also
 give their atomic
 number. Chapter 4

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 (i) Atomic number =
 number of protons +
 number of electrons (ii)
 Mass number =
 number of protons +
 number of neutrons (iii)
 Atomic mass = number
 of protons = number of
 neutrons (iv) Atomic
 number = number of
 protons = number of
 electrons (a) (i) and (ii)
 (b) (i) and (iii) (c) (ii)
 and (iii) (d) (ii) and (iv)

Answer. Answer: (d) (ii)
 and (iv)
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 the sphere =
 $\left(\frac{4}{3}\right) \pi R^3$ Let
 R be the radius of the
 atom and r be that of
 the nucleus. Question
 6. Show
 diagrammatically the

electron distribution in
 a sodium atom and a
 sodium ion and also
 give their atomic
 number.

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atoms be seen
answer the challenges
to his ideas,
Democritus's atomic
theory was eventu-
ment, a narrow beam
of alpha particles was
aimed at a thin sheet
of gold foil.
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Structure of the
Atom - Mrs.Taylor's**

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Questions Circle The
Best Answer 1 Which
Of The Following Are
Isotopes Of Each Other
A 14c And 14n B 3h
And 4he C 2h And 1h D
None Of These 2 The
Net Charge On An
Atom That Has 13
Protons 12 Neutrons
And 10 Electrons Is A 5
B 3 C 3 D 9 3 Particles
That Make Up Atoms
Are Called Subatomic
Particles And Are
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and faster can be
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ways. Experiencing,

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The sum of the protons and neutrons in the nucleus of an atom.

$MASS\# = \#$
protons(atomic number) + #neutrons.
 $\#NEUTRONS = Mass \# - \#Protons(atomic\ number)$
 $\#Protons = Mass \# - \#Neutrons.$
ISOTOPES: Are atoms of the same element that have different numbers of. neutrons and different mass numbers.

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 and 6 neutrons. The
 total number of
 electrons present in an
 atom of this element is
 8 What is the
 maximum number of
 electrons that can
 occupy the second
 energy level?
 study of the structure
 of the atom. STEP 1
 Fold a sheet of paper in
 half lengthwise. Make
 the back edge about 2
 cm longer than the
 front edge. Step 2 Fold
 into thirds. Step 3
 Unfold and cut along
 one fold line to make
 one small tab and one
 large tab. Step 4 Label
 as shown. &/,,\$!",
 Use this Foldable with
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